

1. Define the following:

- a) Covalent Bond _____
- b) Ionic Bond _____
- c) Cation _____
- d) Anion _____

2. Complete the following table:

Ionic Compound	Cation(s)	Anion(s)
<i>Example:</i> $Al_2(CO_3)_3$	$2 Al^{3+}$	$3 CO_3^{2-}$
K_2S		
$CoSO_4$		
$Ni_3(PO_4)_2$		
SnO		
$(NH_4)_3PO_4$		

3. Write the formula of the ionic compound formed between each pair.

Example: Ba^{2+} and Cl^- $BaCl_2$

- a) Al^{3+} and S^{2-} _____
- b) Be^{2+} and F^- _____
- c) Ba^{2+} and O^{2-} _____
- c) K^+ and N^{3-} _____

4. Write the name for each of the following compounds:

SO ₃	
KBr	
HNO ₃	
Cu ₂ O	
Fe ₂ (SO ₄) ₃	
N ₂ O ₄	
K ₃ PO ₄	
Al ₂ O ₃	
NH ₄ Cl	
H ₂ SO ₄	
SeO ₃	
Be(ClO ₄) ₂	
(NH ₄) ₂ Cr ₂ O ₇	
Ba(BrO ₃) ₂	
XeF ₂	
Al ₂ S ₃	
Na ₂ HPO ₄	
Mg ₃ (PO ₄) ₂	
Al(OH) ₃	
CuSO ₃	
Li ₂ HPO ₄	
Ca(NO ₃) ₂	
Cr ₂ (SO ₃) ₃	
Ni(ClO ₄) ₂	
HClO	

5. Write the formula for each of the following compounds:

barium nitrate	
phosphoric acid	
potassium iodide	
iron (III) bromide	
copper (I) phosphate	
zinc sulfite	
calcium hydroxide	
tin (IV) sulfide	
iron (II) chloride	
sulfurous acid	
sodium acetate	
ferric bicarbonate	
zinc sulfite	
silver bicarbonate	
potassium iodide	
barium bisulfite	
lead(IV) chlorite	
nitric acid	
calcium sulfide	
lead(II) nitrite	
copper(I) bisulfate	
potassium dichromate	
sulfuric acid	
boron phosphide	
cobaltic chlorate	