

CHM 100 REVIEW FOR TEST 3: CHAP. 9, 4, 13, 16  
50 MULTIPLE CHOICE QUESTIONS 100 POINTS

TOPICS	M.C.
<b>CHAPTER 9</b> REACTION STOICHIOMETRY: MOLE-MOLE, MASS-MASS PERCENT YIELD PROBLEMS LIMITING REACTANT PROBLEMS	<b>20</b>
<b>CHAPTERS 4 &amp; 13</b> IDENTIFYING AND CONVERSIONS OF UNITS PRESSURE, TEMPERATURE, VOLUME, MOLES CONDITIONS OF STP COMBINED GAS LAWS BOYLE'S, CHARLES', GAY-LUSSAC'S, COMBINED IDEAL GAS LAW MOLAR MASS & DENSITY	<b>20</b>
<b>CHAPTERS 16</b> DEFINING AQUEOUS SOLUTIONS IONIC COMPOUNDS DISSOLVED IN WATER COVALENT COMPOUNDS DISSOLVED IN WATER SOLUTION TERMINOLOGY SOLUBILITY CONCENTRATION UNSATURATED, SATURATED, SUPERSATURATED MOLARITY	<b>9</b>
<b>BONUS QUESTION</b>	<b>1</b>

**SUGGESTIONS FOR STUDY:**

If possible, find someone in the class to study with or make an appointment with your group to meet in the library to study. Studying with another person will reduce the amount of time needed to learn the material

- Review the lecture notes with FlexText Information.
- PRACTICE, PRACTICE, PRACTICE Problems.
- Write out the steps you use for solving problems.
- Familiarize yourself with terms and definitions.
- Practice group problems and suggested textbook problems (at top of the lecture notes).
- Practice the example test on the website. Use the periodic table and provided equations you will be using on the test.

Information provided:

1 mole = $6.02 \times 10^{23}$ At STP, 1 mol of any gas = 22.4 L	$^{\circ}\text{F} = 1.8^{\circ}\text{C} + 32, \text{K} = 273.15 + ^{\circ}\text{C}$
1 atm = 760. torr = 760. mmHg 1 atm = 101.3 kPa	$M_m = \frac{mRT}{PV} \quad M_m = \frac{dRT}{P} \quad d = \frac{M_m P}{RT}$
$\frac{P_1 V_1}{n_1 T_1} = \frac{P_2 V_2}{n_2 T_2}$ $P_1 V_1 T_2 = P_2 V_2 T_1$	$PV = nRT \quad R = 0.0821 \frac{\text{atmL}}{\text{molK}}$
<b>Molarity = moles solute/volume solution</b>	$M_1 V_1 = M_2 V_2$