

TOPICS	M.C.
UNIT 1: CHEMISTRY & MATTER ARISTOTLE THE PHILOSOPHER'S STONE STATES OF MATTER PHYSICAL AND CHEMICAL PROPERTIES AND CHANGES PHLOGISTON CLASSIFYING MATERIALS HETEROGENEOUS MATERIALS HOMOGENEOUS MATERIALS SOLUTIONS PURE SUBSTANCES ELEMENTS COMPOUNDS CHARACTERISTICS OF A CHEMICAL CHANGE CHEMICAL FORMULAS, CHEMICAL EQUATIONS ENERGY DEFINITIONS	22
UNIT 2: THE METRIC SYSTEM – KNOW THE CONVERSIONS! SI BASE UNITS UNIT CONVERSION PROBLEMS SCIENTIFIC NOTATION DIFFERENT MEASUREMENTS: WEIGHT & MASS, LENGTH, TEMPERATURE DERIVED UNITS: VOLUME & DENSITY PRECISION VS. ACCURACY SIGNIFICANT FIGURES	23
UNIT 3: LAW OF CONSERVATION OF MASS SUBATOMIC PARTICLES STRUCTURE OF THE ATOM STRUCTURE OF THE NUCLEUS CATION, ANION ATOMIC NUMBER MASS NUMBER ISOTOPES	5

SUGGESTIONS FOR STUDY:

If possible, find someone in the class to study with or make an appointment with your group to meet in the library to study. Studying with another person will reduce the amount of time needed to learn the material.

- Review the lecture notes with Textbook Information.
- Write out the methods used in calculations and the steps to follow.
- Write out flash cards for items you must memorize.
- Familiarize yourself with the tables you will be using.
- Practice group problems and suggested textbook problems (at top of the lecture notes).
- Practice the example test on the website.

INFORMATION PROVIDED:

$$D = m/V \quad ^\circ\text{F} = 1.8^\circ\text{C} + 32 \quad ^\circ\text{C} = (^\circ\text{F} - 32)/1.8 \quad \text{K} = 273.15 + ^\circ\text{C}$$