

Objectives for Unit 6:

- ü To be able to write and balance an unbalanced chemical equation.
- ü To be able to write balanced equations for combination, decomposition, oxidation - reduction, and double replacement reactions.
- ü To use solubility rules to help with some general rules for water solubility's of common ionic compounds.

EVOLUTION OF A CHEMICAL EQUATION

TERMS AND SYMBOLS

BALANCING CHEMICAL EQUATIONS

CLASSIFYING CHEMICAL REACTIONS

1. COMBINATION (SYNTHESIS) REACTIONS
2. DECOMPOSITION REACTIONS
3. COMBUSTION REACTIONS
4. SINGLE REPLACEMENT REACTIONS (OXIDATION - REDUCTION REACTIONS)
5. DOUBLE REPLACEMENT REACTIONS
 - a. PRECIPITATE REACTIONS (STILL DOUBLE REPLACEMENT REACTIONS)
 - PRECIPITATE
 - SOLUBLE
 - INSOLUBLE
 - SOLUBILITY RULES
 - b. NEUTRALIZATION REACTIONS (STILL DOUBLE REPLACEMENT)